

## Chapter 12 Physical Properties Of Solutions

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**Chapter 12 Physical Properties Of**  
Chapter 12. Chapter 12: Physical Properties of Solutions ===== [12.1]. Which of the following liquids would make a good solvent for bromine, Br 2? A. CCl 4 B. H 2 O C. CH 3 OH D. NH 3 [A] ----- [12.2]. A 20.0 % by mass solution of phosphoric acid (H3PO4) in water has a density of 1.114 g/mL at 20°C. What is the molarity of this solution?

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234 12: Physical Properties of Solutions ~- AHsoln = AH- + AH2 + AH3 The solute will be soluble in the solvent if the solut~r c ~on is stronger than s'luti rt\_p[access is , ~n is weaker th~-- solution ~-ic, Sol~--~N= a solute \* ~-, (an loa er molecule) i& ~urround~d b~--o str6~e-solvent a~ractive forces.

**Chapter 12. Physical Properties of Solutions**  
Chapter 12 | Physical Properties of Solutions. Students will be able to: • Identify different types of solutions. • Understand the solvation process at a molecular level. • Calculate the...

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Chapter 12: Physical Properties of Solutions 1. A saturated solution A) contains more solute than solvent. B) contains more solvent than solute. C) contains equal moles of solute and solvent. D) contains the maximum amount of solute that will dissolve in that solvent at that temperature.

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As shown in Table 12.5 "Physical Properties of Some Alkanes", the boiling points of the straight-chain alkanes increase with increasing molar mass.This general rule holds true for the straight-chain homologs of all organic compound families. Larger molecules have greater surface areas and consequently interact more strongly; more energy is therefore required to separate them.

**12.6 Physical Properties of Alkanes | The Basics of ...**  
CHAPTER 12: PHYSICAL PROPERTIES OF SOLUTIONS 314 12.15 Percent mass equals the mass of solute divided by the mass of the solution (that is, solute plus solvent) times 100 (to convert to percentage).

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View Notes - Chapter 12 Lecture 1 from CHEM 1123 at University of Arkansas. Chapter 12: Physical Properties of Solutions Solution = Solvent + Solute Homogeneous Examples: CH4 + H2S CO2 +